



La₃B₁₄[−]: an inverse triple-decker lanthanide boron cluster†

Teng-Teng Chen,^{‡a} Wan-Lu Li,^{‡b} Wei-Jia Chen,^{‡a} Jun Li^{‡*bc} and Lai-Sheng Wang^{‡*a}

Cite this: *Chem. Commun.*, 2019, 55, 7864

Received 16th May 2019,
Accepted 13th June 2019

DOI: 10.1039/c9cc03807h

rsc.li/chemcomm

We report the observation of the first inverse triple-decker complex in a tri-lanthanide-doped boron cluster. Photoelectron spectroscopy of La₃B₁₄[−] reveals well-resolved photodetachment transitions. Quantum chemical studies show that the most stable structure of the La₃B₁₄[−] cluster exhibits a tilted La–B₈–La–B₈–La inverse triple-decker structure with two conjoined B₈ rings sharing a pair of B atoms due to strong inter-layer B–B bonding. The tilted structure enhances both B–B and B–La bonding, resulting in a highly stable inverse triple-decker structure. Theoretical calculations further show that multi-decker conjoined structures are viable as a new class of 1D lanthanide boron nanostructures.

Boron has many allotropes with complex crystal structures and super-hard properties because of its electron deficiency and the strong covalent B–B bonding.^{1,2} Boron can also form a variety of metal boride materials with important potential industrial applications, ranging from super-hard metal borides and superconducting MgB₂ to boride materials with ultra-high magnetic fields and ultra-high thermal conductivity.^{3–5} Nanoclusters of boron have also been shown to display a diverse range of interesting stable structures and multi-center chemical bonding.^{6–10} Joint experimental and theoretical studies in the past decade on size-selected boron clusters have revealed planar (2D) structures leading to the discovery of graphene-like planar structures (borophenes), fullerene-like cage structures (borospherenes), and nanotubular structures.^{6–13} Metal-doped boron clusters have led to a new direction in the study of boron nanoclusters, resulting in the discovery of metal-centered

molecular wheels, half-sandwich structures, possibilities of metallo-borophenes, and nanotubular drum-like structures.^{14–19} Most recently, a new class of inverse-sandwich lanthanide boride clusters containing monocyclic and aromatic B_n rings (M–B_n–M, *n* = 7–9) have been uncovered *via* joint photoelectron spectroscopy (PES) and quantum chemistry calculations.^{20,21} These exciting findings immediately lead to the question of whether triple-decker or even multi-decker lanthanide boride clusters would be possible.

Triple-decker complexes (L–M–L–M–L, L = aromatic ligands) represent a class of interesting organometallic compounds important in organic and inorganic syntheses, as well as building blocks for molecular architectures.^{22–25} The first triple-decker transition-metal complex was synthesized in 1972,²⁶ followed by electronic structure studies and the syntheses of a series of new triple-decker compounds.^{27–30} In the past four decades, a variety of triple-decker complexes have been synthesized with different aromatic ligands and d-/f-metal atoms, which can be potential one-dimensional (1D) magnetic materials and molecular information storage media.^{25,31–34} However, inverse triple-decker (M–L–M–L–M) or multi-decker lanthanide complexes have not been synthesized.

Here we report the first observation and characterization of a tri-lanthanum-doped boron cluster, La₃B₁₄[−], which is found to possess a tilted inverse triple-decker structure using a joint PES and theoretical investigation. Global minimum searches have shown that the structure can be viewed as two conjoined B₈ rings sharing two B atoms with C_{2v} symmetry. Theoretical calculations show that multi-decker La_nB_{6n–4}[−] systems with conjoint B₈ rings can form stable 1D lanthanide–boron nano-wires. Such novel 1D nanostructures may be viable to be realized experimentally on suitable substrates or found as motifs in new lanthanide borides.

The La₃B₁₄[−] cluster was produced using laser-vaporization of a mixed La/¹¹B disk target and characterized using a magnetic-bottle PES apparatus (see the ESI† for details).^{10,35} The photoelectron spectrum of La₃B₁₄[−] at 193 nm is shown in Fig. 1a, which is well resolved with distinct spectral features,

^a Department of Chemistry, Brown University, Providence, Rhode Island 02912, USA. E-mail: Lai-Sheng_Wang@brown.edu

^b Department of Chemistry and Key Laboratory of Organic Optoelectronics & Molecular Engineering of Ministry of Education, Tsinghua University, Beijing 100084, China. E-mail: junli@tsinghua.edu.cn

^c Department of Chemistry, Southern University of Science and Technology, Shenzhen, Guangdong 518055, China

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9cc03807h

‡ These authors contributed equally to this work.

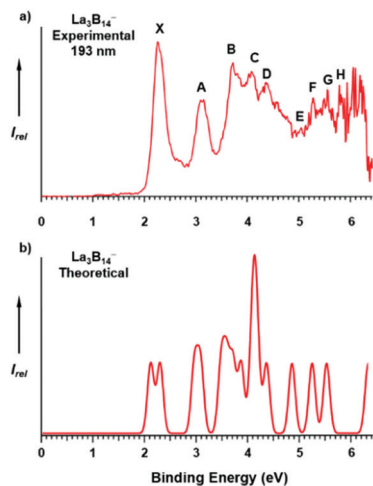


Fig. 1 (a) Photoelectron spectrum of $\text{La}_3\text{B}_{14}^-$ at 193 nm and (b) the simulated spectrum of the tilted triple-decker global minimum of $\text{La}_3\text{B}_{14}^-$ at the SAOP/TZP level.

labeled as X and A–H. The X band corresponds to the detachment transition from the ground state of $\text{La}_3\text{B}_{14}^-$ to that of neutral La_3B_{14} , while the bands at higher binding energies come from detachment transitions to excited states of neutral La_3B_{14} . The overall spectral pattern is relatively simple and well-resolved for such a complicated binary cluster, suggesting a highly stable and symmetric global minimum. Band X is sharp and intense, giving rise to the first vertical detachment energy (VDE) of 2.25 eV for $\text{La}_3\text{B}_{14}^-$. The adiabatic detachment energy (ADE) of band X was estimated from its onset to be 2.05 eV, which also represents the electron affinity (EA) of neutral La_3B_{14} . Following an energy gap, band A at 3.15 eV is well separated from other spectral features. Following another energy gap from band A, a congested spectral range ensues from about 3.5 to 4.5 eV. Three bands are labeled, B, C, and D, which likely contain multiple detachment transitions. Between 4.5 and 6.2 eV, again congested spectral features are observed and four bands, E, F, G, and H, are tentatively identified. The binding energies of all detachment features are given in Table S1 (ESI†). The PES spectral features and the overall spectral pattern of $\text{La}_3\text{B}_{14}^-$ can serve as an electronic fingerprint that can be used to compare with theoretical calculations to probe its structure, stability, and chemical bonding.

The global minimum of $\text{La}_3\text{B}_{14}^-$ was searched using the TGMMin 2.0 package^{36,37} combined with calculations employing the ADF software.³⁸ We examined more than 1000 structures and found that a tilted inverse triple-decker structure with C_{2v} symmetry was the global minimum on the potential energy surface, as shown in Fig. 2, and the Cartesian coordinates are listed in Table S2 (ESI†). The C_{2v} structure can be viewed as the fusion of two B_8 monocyclic rings sharing two B atoms. The structural relation between the C_{2v} global minimum and a perfect $\text{La-B}_8\text{-La-B}_8\text{-La}$ triple decker will be discussed later (*vide infra*). Low-lying isomers within 54 kcal mol⁻¹ of the global minimum are given in Fig. S1 (ESI†). The second lowest energy isomer (C_1) is found to be 10.92 kcal mol⁻¹ higher in

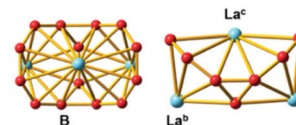


Fig. 2 The global minimum structure of $\text{La}_3\text{B}_{14}^-$ (C_{2v} , 1A_1) with two conjoined B_8 rings: top view (left) and side view (right). La^b and La^c represent the outer and central La atoms, respectively.

energy than the C_{2v} global minimum at the PBE0/TZP level, highlighting the stability of the tilted triple-decker structure. We also calculated the linear $\text{La-B}_7\text{-La-B}_7\text{-La}$ triple-decker structure with two B_7 monocyclic rings and found two imaginary frequencies involving the bending modes (tilting) of the two B_7 rings. Furthermore, the energy of this triple-decker structure lies 150 kcal mol⁻¹ above the C_{2v} global minimum. Following the imaginary frequencies would lead to the global minimum.

To validate the global minimum of the $\text{La}_3\text{B}_{14}^-$ cluster, we calculated the first ADE and VDEs of the C_{2v} tilted triple-decker structure using the PBE and PBE0 methods with the TZP basis sets. All the calculated detachment channels together with the electron configurations and final state symmetries are compared with the experimental data in Table S1 (ESI†). Each VDE is fitted with a unit-area Gaussian of 0.15 eV width to produce a simulated spectrum, as shown in Fig. 1b. The computed first ADE/VDE values are 2.12/2.18 eV, in good agreement with the experimental results of 2.05/2.25 eV.

The C_{2v} tilted triple-decker structure of $\text{La}_3\text{B}_{14}^-$ has a closed-shell electron configuration and its valence molecular orbitals (MOs) are depicted in Fig. S2 (ESI†). Hence, electron detachment from each occupied MO results in one doublet final state (Table S1, ESI†), yielding a relatively simple photoelectron spectrum. Detachment from the $24b_2$ HOMO gives rise to a computed VDE of 2.18 eV, whereas detachment from the $18a_2$ HOMO–1 results in a computed VDE of 2.36 eV. These two detachment channels have closely spaced binding energies, consistent with the intense band X. Detachments from the $45a_1$ and $17a_2$ MOs also give rise to two closely spaced VDEs of 3.05 eV and 3.17 eV, respectively, which are well separated from the first two detachment channels, in excellent agreement with band A (an experimental VDE of 3.15 eV). Following an energy gap and starting from the $34b_1$ MO, which gives a computed VDE of 3.58 eV, a series of closely spaced detachment channels are computed up to 5.7 eV, consistent with the congested PE spectral features beyond band A (Fig. 1 and Table S1, ESI†). Overall, the simulated spectrum from the global minimum of $\text{La}_3\text{B}_{14}^-$ agrees well with the experimental data, providing considerable credence for the identified tilted triple-decker structure.

The chemical bonding of the tilted triple-decker global minimum of $\text{La}_3\text{B}_{14}^-$ is analyzed using the adaptive natural density partitioning (AdNDP) method.³⁹ The AdNDP results reveal that the 52 valence electrons in $\text{La}_3\text{B}_{14}^-$ can be classified into six categories, as displayed in Fig. 3. Eleven two-center two-electron (2c–2e) σ bonds are found, five on each B_8 ring plus one between the two shared boron atoms. There are also two 3c–2e delocalized bonds on the B_3 units shared by the two B_8

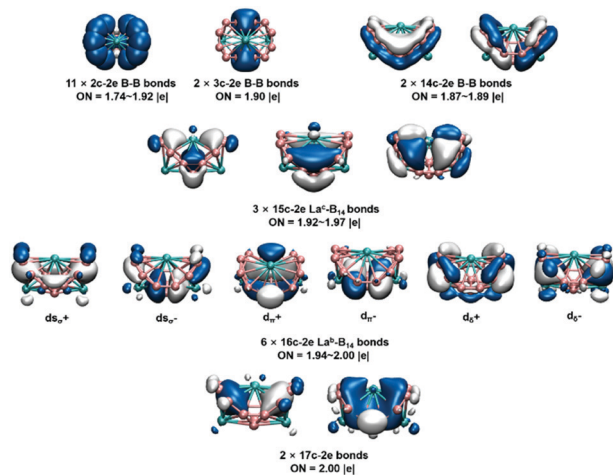


Fig. 3 AdNDP analyses of the chemical bonding in the tilted inverse triple-decker global minimum of $\text{La}_3\text{B}_{14}^-$ at the PBE0/TZP level. The occupation number (ON) is given.

rings (Fig. 2). There are two 14c-2e bonds with π characters representing delocalized bonding within the two conjoined B_8 rings with little bonding with the La atoms. The remaining 11 multicenter bonds describe various levels of interactions between the conjoined B_8 ligands and the three La atoms in the tilted triple-decker. The three delocalized 15c-2e bonds in the second row represent mainly interactions between the central La atom (labeled as La^c) and the boron framework, involving the La^c 5d π and 5d δ orbitals. The two outer La atoms (labeled as La^b) interact with the fourteen B atoms *via* six 16c-2e bonds, as depicted in the third row. Different interactions are defined according to the z-axis along the $\text{La}^b \cdots \text{La}^b$ direction: “+” and “-” represent positive and negative overlaps, respectively. Finally, the two 17c-2e bonds in the fourth row describe bonding between the three La atoms and the boron framework in the tilted triple-decker, with primary interactions involving the two outer La^b atoms and the boron framework with some contribution from the central La^c atom.

Recently, the first inverse sandwich consisting of a B_8 ring and two La atoms (D_{8h} La_2B_8) has been observed. The B-B and B-La bond distances are 1.56 and 2.76 Å, respectively. The global minimum of $\text{La}_3\text{B}_{14}^-$ can be viewed as the first inverse triple-decker consisting of two B_8 rings. However, the perfect linear $\text{La} \cdots \text{B}_8 \cdots \text{La} \cdots \text{B}_8 \cdots \text{La}$ inverse triple-decker is not stable because of the strong inter-layer B-B bonding. The transformation from this high-symmetry linear structure to the bent global minimum of $\text{La}_3\text{B}_{14}^-$ is schematically shown in Fig. 4, by tilting the two B_8 rings and sharing one B-B unit. There are some structural distortions in the conjoined part of the two rings, whereas the unshared five B-B bonds on each B_8 ring are very similar to those in the La_2B_8 inverse sandwich, ranging from 1.53 to 1.60 Å (Fig. S3, ESI†). These B-B bond lengths are consistent with B=B double bonds, according to Pyykko's covalent radii.⁴⁰ The multiple B-B bond order is borne out from the AdNDP analyses for $\text{La}_3\text{B}_{14}^-$ (Fig. 3), as well as for La_2B_8 . In the conjoined part, the shared B-B bond has a

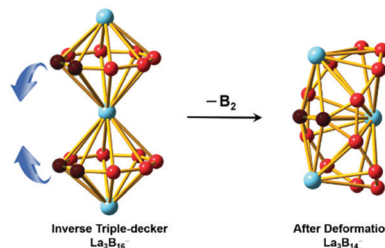


Fig. 4 Schematics showing the transformation from a perfect linear inverse triple-decker, $\text{La} \cdots \text{B}_8 \cdots \text{La} \cdots \text{B}_8 \cdots \text{La}$, to the global minimum, bent triple-decker of $\text{La}_3\text{B}_{14}^-$ with two conjoined B_8 rings by sharing a B_2 unit (dark-brown color).

significantly lengthened bond distance of 1.78 Å, which is consistent with a B-B single bond. The neighboring B atoms next to the shared B-B bond from both the B_8 rings also form a B-B single bond (1.70 Å, Fig. S3, ESI†), resulting in a B_3 triangle with one of the shared B atoms. The 3c-2e bonds revealed in the AdNDP analyses are largely responsible for these additional conjoined B-B bonds, which significantly enhanced the binding between the two conjoined B_8 rings.

The conjoined B_8 rings did not affect the bonding with the La atoms, in fact, they have enhanced the metal-ligand bonding. The La^b -B bond distances for the two outer La atoms range from 2.68 to 2.78 Å, which are shorter on average than those in the La_2B_8 inverse sandwich (2.76 Å). The central La atom (La^c) also has optimal bonding with the boron framework, as shown in the AdNDP analyses discussed above. We also carried out QTAIM calculation and analyzed bond order indices to investigate the bonding of La-B and the bridged B-B unit as shown in Table S3 (ESI†). Compared with covalent B-B interactions, the La-B bonding is primarily of dative bonding character ($\rho > 0$, $E < 0$, $\nabla^2\rho > 0$) with little covalent contribution. The binding energy between the three La atoms and the boron framework is calculated to be 665.15 kcal mol⁻¹ at the PBE0/TZP level. Overall, the high stability of the global minimum of $\text{La}_3\text{B}_{14}^-$ is derived from both the enhanced B-B and La-B bonding, making it an exceptional inverse triple-decker complex.

For transition-metal-doped boron clusters, previous studies have shown metal-centered tubular structures,^{16,17,19,41,42} such as $\text{D}_{8d}\text{-Co}(\text{B}_8)_2^-$, $\text{D}_{9d}\text{-Rh}(\text{B}_9)_2^-$, and $\text{D}_{10d}\text{-Ta}(\text{B}_{10})_2^-$, which suggested the possibilities of metal-centered 1D boron nanotubes. The high stability of the $\text{La}_3\text{B}_{14}^-$ tilted inverse triple-decker immediately points to a new type of 1D lanthanide-boron nanostructure by extending the triple decker with conjoint B_8 rings. We have carried out preliminary calculations for two members of this series, $\text{La}_5\text{B}_{26}^-$ consisting of four conjoint B_8 rings and $\text{La}_7\text{B}_{38}^-$ consisting of six conjoint B_8 rings, as shown in Fig. 5. The optimized structure of $\text{La}_5\text{B}_{26}^-$ is closed-shell with C_{2v} (1A_1) symmetry and that of $\text{La}_7\text{B}_{38}^-$ is closed-shell with C_s ($^1A'$) symmetry. These 1D oligomeric nanoclusters are true minima on the potential energy surfaces without imaginary vibrational frequencies and their Cartesian coordinates are given in Table S4 (ESI†). There should be a whole series of such 1D nanoclusters with the chemical formula of $\text{La}_n\text{B}_{6n-4}^-$. Even an infinitely long 1D nanowire may be viable or exist in new

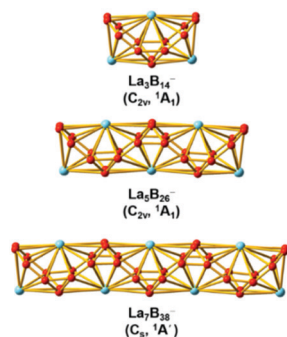


Fig. 5 Optimized structures of $\text{La}_5\text{B}_{26}^-$ and $\text{La}_7\text{B}_{38}^-$ compared with that of $\text{La}_3\text{B}_{14}^-$ at the SR-ZORA PBE/TZP level.

lanthanide boride compounds. Considering the fact that most of the lanthanide (Ln = La–Lu) elements can form similar inverse boron sandwiches,²⁰ we expect that they can also form similar tilted triple-deckers and 1D nanowires with tunable magnetic properties of multiple $4f^n$ atomic centers.

In summary, we have observed the first inverse lanthanide–boron triple-decker cluster. Photoelectron spectroscopy of $\text{La}_3\text{B}_{14}^-$ revealed a relatively simple spectrum, suggesting a symmetric and stable structure. Theoretical calculations have shown that the global minimum of $\text{La}_3\text{B}_{14}^-$ is closed-shell with C_{2v} symmetry, which can be viewed as a tilted inverse triple-decker with two conjoined B_8 rings sharing a B–B unit due to strong inter-layer B–B bonding. The features of the B_8 rings can still be recognized despite the strong inter-ring interactions. The conjoined B_8 rings also enhance interactions with the La atoms, resulting in the high stability of the tilted inverse triple-decker. Extension of the tilted triple-decker can result in a series of 1D nanoclusters of the formula $\text{La}_n\text{B}_{6n-4}^-$. Infinitely long 1D lanthanide–boron nanostructures consisting of conjoined B_8 rings are viable with different magnetic properties or may be found in crystalline lattices of new lanthanide boride materials.

The experimental work done at Brown University was supported by the U.S. National Science Foundation (CHE-1763380). The theoretical work done at Tsinghua University was supported by the National Natural Science Foundation of China (Grant No. 21590792, 91426302, and 21433005). The calculations were performed using resources from the Computational Chemistry Laboratory of the Department of Chemistry under the Tsinghua Xuetang Talents Program.

Conflicts of interest

There are no conflicts to declare.

Notes and references

- 1 A. R. Oganov, J. Chen, C. Gatti, Y. Ma, Y. Ma, C. W. Glass, Z. Liu, T. Yu, O. O. Kurakevych and V. L. Solozhenko, *Nature*, 2009, **457**, 863–867.
- 2 B. Albert and H. Hillebrecht, *Angew. Chem., Int. Ed.*, 2009, **48**, 8640–8668.
- 3 J. Nagamatsu, N. Nakagawa, T. Muranaka, Y. Zenitani and J. Akimitsu, *Nature*, 2001, **410**, 63–64.
- 4 H. Y. Chung, M. B. Weinberger, J. B. Levine, A. Kavner, J. M. Yang, S. H. Tolbert and R. B. Kaner, *Science*, 2007, **316**, 436–439.
- 5 F. Tian and Z. Ren, *Angew. Chem., Int. Ed.*, 2019, **58**, 5824–5831.
- 6 H. J. Zhai, B. Kiran, J. Li and L. S. Wang, *Nat. Mater.*, 2003, **2**, 827.
- 7 A. N. Alexandrova, A. I. Boldyrev, H. J. Zhai and L. S. Wang, *Coord. Chem. Rev.*, 2006, **250**, 2811–2866.
- 8 E. Oger, N. R. M. Crawford, R. Kelting, P. Weis, M. M. Kappes and R. Ahlrichs, *Angew. Chem., Int. Ed.*, 2007, **46**, 8503–8506.
- 9 A. P. Sergeeva, I. A. Popov, Z. A. Piazza, W. L. Li, C. Romanescu, L. S. Wang and A. I. Boldyrev, *Acc. Chem. Res.*, 2014, **47**, 1349–1358.
- 10 L. S. Wang, *Int. Rev. Phys. Chem.*, 2016, **35**, 69–142.
- 11 Z. A. Piazza, H. S. Hu, W. L. Li, Y. F. Zhao, J. Li and L. S. Wang, *Nat. Commun.*, 2014, **5**, 3113.
- 12 H. J. Zhai, Y. F. Zhao, W. L. Li, Q. Chen, H. Bai, H. S. Hu, Z. A. Piazza, W. J. Tian, H. G. Lu, Y. B. Wu, Y. W. Mu, G. F. Wei, Z. P. Liu, J. Li, S. D. Li and L. S. Wang, *Nat. Chem.*, 2014, **6**, 727.
- 13 B. Kiran, S. Bulusu, H. J. Zhai, S. Yoo, X. C. Zeng and L. S. Wang, *Proc. Natl. Acad. Sci. U. S. A.*, 2005, **102**, 961–964.
- 14 C. Romanescu, T. R. Galeev, W. L. Li, A. I. Boldyrev and L. S. Wang, *Acc. Chem. Res.*, 2012, **46**, 350–358.
- 15 I. A. Popov, W. L. Li, Z. A. Piazza, A. I. Boldyrev and L. S. Wang, *J. Phys. Chem. A*, 2014, **118**, 8098–8105.
- 16 I. A. Popov, T. Jian, G. V. Lopez, A. I. Boldyrev and L. S. Wang, *Nat. Commun.*, 2015, **6**, 8654.
- 17 W. L. Li, X. Chen, T. Jian, T. T. Chen, J. Li and L. S. Wang, *Nat. Rev. Chem.*, 2017, **1**, 0071.
- 18 T. T. Chen, W. L. Li, T. Jian, X. Chen, J. Li and L. S. Wang, *Angew. Chem., Int. Ed.*, 2017, **56**, 6916–6920.
- 19 W. L. Li, T. Jian, X. Chen, H. R. Li, T. T. Chen, X. M. Luo, S. D. Li, J. Li and L. S. Wang, *Chem. Commun.*, 2017, **53**, 1587–1590.
- 20 W. L. Li, T. T. Chen, D. H. Xing, X. Chen, J. Li and L. S. Wang, *Proc. Natl. Acad. Sci. U. S. A.*, 2018, **115**, E6972–E6977.
- 21 T. T. Chen, W. L. Li, J. Li and L. S. Wang, *Chem. Sci.*, 2019, **10**, 2534–2542.
- 22 R. Hoffmann, *Angew. Chem., Int. Ed. Engl.*, 1982, **21**, 711–724.
- 23 A. Fessenbecker, M. D. Attwood, R. F. Bryan, R. N. Grimes, M. K. Woode, M. Stephan, U. Zenneck and W. Siebert, *Inorg. Chem.*, 1990, **29**, 5157–5163.
- 24 D. K. P. Ng and J. Jiang, *Chem. Soc. Rev.*, 1997, **26**, 433–442.
- 25 V. Lorenz, S. Blaurock, C. G. Hrib and F. T. Edelmann, *Organometallics*, 2010, **29**, 4787–4789.
- 26 A. Salzer and H. Werner, *Angew. Chem., Int. Ed. Engl.*, 1972, **11**, 930–932.
- 27 J. W. Lauher, M. Elian, R. H. Summerville and R. Hoffmann, *J. Am. Chem. Soc.*, 1976, **98**, 3219–3224.
- 28 H. Werner, *Angew. Chem., Int. Ed. Engl.*, 1977, **16**, 1–9.
- 29 J. Moraczewski and W. E. Geiger Jr, *J. Am. Chem. Soc.*, 1978, **100**, 7429–7431.
- 30 A. R. Kudinov, M. I. Rybinskaya, Y. T. Struchkov, A. I. Yanovskii and P. V. Petrovskii, *J. Organomet. Chem.*, 1987, **336**, 187–197.
- 31 W. Siebert, J. Edwin and M. Bochmann, *Angew. Chem., Int. Ed. Engl.*, 1978, **17**, 868–869.
- 32 E. D. Jemmis and A. C. Reddy, *Organometallics*, 1988, **7**, 1561–1564.
- 33 D. Chabach, A. De Cian, J. Fischer, R. Weiss and M. E. M. Bibout, *Angew. Chem., Int. Ed. Engl.*, 1996, **35**, 898–899.
- 34 J. Li, D. Gryko, R. B. Dabke, J. R. Diers, D. F. Bocian, W. G. Kuhr and J. S. Lindsey, *J. Org. Chem.*, 2000, **65**, 7379–7390.
- 35 L. S. Wang, H. S. Cheng and J. Fan, *J. Chem. Phys.*, 1995, **102**, 9480–9493.
- 36 Y. F. Zhao, X. Chen and J. Li, *Nano Res.*, 2017, **10**, 3407–3420.
- 37 X. Chen, Y. F. Zhao, Y. Y. Zhang and J. Li, *J. Comput. Chem.*, 2019, **40**, 1105–1112.
- 38 ADF, 2016.101, SCM, Theoretical Chemistry, Vrije Universiteit, Amsterdam, The Netherlands, (<http://www.scm.com>).
- 39 D. Y. Zubarev and A. I. Boldyrev, *Phys. Chem. Chem. Phys.*, 2008, **10**, 5207–5217.
- 40 P. Pyykkö, *J. Phys. Chem. A*, 2015, **119**, 2326–2337.
- 41 T. Jian, W. L. Li, I. A. Popov, G. V. Lopez, X. Chen, A. I. Boldyrev, J. Li and L. S. Wang, *J. Chem. Phys.*, 2016, **144**, 154310.
- 42 T. Jian, W. L. Li, X. Chen, T. T. Chen, G. V. Lopez, J. Li and L. S. Wang, *Chem. Sci.*, 2016, **7**, 7020–7027.