



ELSEVIER

Contents lists available at ScienceDirect

## Fire Safety Journal

journal homepage: [www.elsevier.com/locate/firesaf](http://www.elsevier.com/locate/firesaf)

## Short Communication

## A simple method to assess the quenching effectiveness of fire suppressants

C.T. Avedisian<sup>a</sup>, C. Presser<sup>b,\*</sup>, J. Stiehl<sup>a</sup>, R.E. Cavicchi<sup>b</sup><sup>a</sup> Sibley School of Mechanical and Aerospace Engineering, Cornell University, Ithaca, NY 14853-7501, USA<sup>b</sup> Chemical Science and Technology Laboratory, National Institute of Standards and Technology, 100 Bureau Dr., Stop 8360, Gaithersburg, MD 20899-8360, USA

## ARTICLE INFO

## Article history:

Received 19 January 2010

Accepted 28 February 2010

## Keywords:

Fire suppressant

Pulse heating

Extinction

Transient heating

Thermal conductivity

Quenching effectiveness

Suppression effectiveness

## ABSTRACT

A screening concept is suggested for evaluating the effectiveness of fluids to thermally suppress fires. It is based on measuring a fluid's ability to inhibit (or quench) the temperature rise of a material that is rapidly heated. The experimental design is similar to the transient hot wire technique, in which the evolution of the average material temperature is recorded for a given input power, and internal temperature gradients in the material are minimized. A gold wire (100 μm long and 5 μm diameter) is used as the surface which heats the fluid. The wire temperature response due to a power pulse provides a measure of the effectiveness of the fluid to suppress thermally the temperature increase. The results indicate that the "quenching effectiveness",  $QE = (T_{\max} - T_{\infty}) / (T_{\max,ref} - T_{\infty})$ , correlates with the ratio of the fluid thermal conductivity to that of the wire,  $k_{fluid}/k_{solid}$ , using different Nusselt numbers (representing both conduction and natural convection) for the liquids or gases. The concept developed here could be included in a more comprehensive screening protocol, which would assess the thermal potential of candidate fire suppressants.

Published by Elsevier Ltd.

## 1. Introduction

A fire suppressant derives its effectiveness by extinguishing a burning surface through chemical, thermal, and physical mechanisms. Previous work [1] is notable for the articulation of thermal, physical, and chemical effects on flame extinction especially for gaseous agents (e.g., trifluoromethane, CF<sub>3</sub>H and bromo-trifluoromethane, CF<sub>3</sub>Br) [2]. The ability to separate suppressant mechanisms when using an actual fire configuration in testing is difficult. Experimental screening methods generally rank agents according to the integrated physical, thermal, and chemical effects that lead to flame extinction [3]. The efficacy of new fire suppressants is facilitated by screening methods that segregate the mechanisms involved in the processes leading to flame extinction. A testing protocol that decouples these effects could provide a clearer understanding of the mechanisms associated with suppressant of a particular fire. Linteris et al. [4] noted that the flame configuration itself could influence the relative suppression effectiveness. At the same time, useful information can be obtained even if the results are unique to that configuration, and indeed, this is the case for standardized screening protocols (e.g., counterflow burner (e.g., [5]), cup burner screens (e.g., [6]), and turbulent spray flames (e.g., [7]).

Several studies have attempted to isolate individual mechanisms of suppression. For example, Pitts et al. [8] ranked suppressants according to their thermophysical properties (i.e., the latent heat of

vaporization, heat capacity, and boiling point were used to estimate the total absorbed heat). Agents were selected according to this ranking to test in various screens for their suppression effectiveness. Lentati and Chelliah [9] used a numerical model to unmask physical, thermal, and chemical effects in a counterflow configuration, in which monodispersed water droplets were introduced into the flow field. A follow-up experimental study [10] indicated that water mists were effective suppressants due to their thermal behavior (as effective as CF<sub>3</sub>Br), and that a strategy of mixing thermal components with chemical components may lead to the development of superior fire suppressants.

Presented is a concept to evaluate the heat transfer properties of fire suppressant fluids, through the interaction of a cold fluid with a hot surface. The experimental approach is based on the pulse heating of a solid material – a metal in the proposed configuration – that is submerged in a test fluid, and monitoring the temperature change of the solid during the power pulse. This configuration is relevant to a re-ignitable hot surface in fire scenarios, as had been investigated by Hamins et al. [3]. We did not consider the condition where vaporization would occur, but the approach is amenable to the characterization of vaporization effects. Vaporization of the fluid would be manifested as an inflection point at the start of nucleation and an increased rate of temperature rise thereafter [11].

## 2. Experimental arrangement

Fig. 1 presents a schematic of the experimental arrangement. In general, a programmed power pulse is imposed on a small

\* Corresponding author. Tel.: +1 301 975 2612.  
E-mail address: [cpresser@nist.gov](mailto:cpresser@nist.gov) (C. Presser).

**Nomenclature**

$Bi$	Biot number ( $=hDk_{solid}^{-1}$ )
$D$	wire diameter (m)
$h$	convection heat transfer coefficient ( $WK^{-1}$ )
$k$	thermal conductivity ( $Wm^{-1}K^{-1}$ )
$L$	wire length (m)
$Nu$	Nusselt number (Eq. (9.34) of Ref. [14])
$p$	power (W)
$\bar{p}$	nondimensional power
$QE$	quenching effectiveness based on the temperature derivative
$R_h$	wire resistance ( $\Omega$ )
$R_{ho}$	wire resistance at room temperature ( $\Omega$ )
$R_p$	adjustable resistor ( $\Omega$ )
$R_1$	bridge resistances ( $\Omega$ )
$t$	time (s)
$T$	wire temperature (K)
$T_{TC}$	oven temperature (K)
$T_{\infty}$	ambient fluid temperature (K)
$V_{in}$	bridge input voltage

$V_{out}$  bridge output voltage (V)

**Greek letters**

$\alpha$	thermal diffusivity ( $m^2s^{-1}$ )
$\Delta t$	time duration (s)
$\theta$	nondimensional temperature difference
$\theta_R$	thermal coefficient of resistivity ( $K^{-1}$ )
$\xi$	suppression effectiveness parameter base on $T_{max}$
$\tau$	nondimensional time

**Subscript**

<i>fluid</i>	referenced to the suppression agent
<i>gas</i>	gaseous agent
<i>liquid</i>	liquid agent
<i>max</i>	referenced to the value near end of pulse (near $t=50\mu s$ )
<i>ref</i>	referenced to the value of water
<i>solid</i>	referenced to the gold wire

diameter metal wire that forms one leg of a bridge circuit. The wire is immersed in the fluid and supported only by its electrical wire-bond connection at either end. During the pulse, the wire electrical resistance is monitored with a fast-transient data acquisition system. The hot wire serves as both an energy source to impart a precise and controllable amount of thermal energy to the surrounding fluid and a temperature sensor through the relationship of electrical resistance with temperature of the wire material from a separate calibration. The fluid that best restrains the temperature change of the wire during the heating process is considered to be the most thermally effective agent.

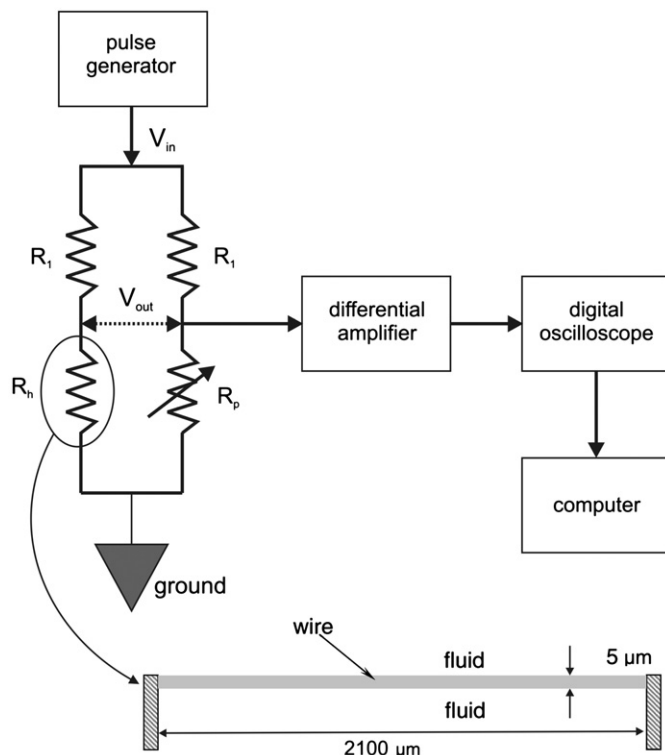


Fig. 1. Schematic of the experimental arrangement.

The gold wire is  $5\mu m$  in diameter ( $D$ ) and  $2.1\text{ mm}$  in length ( $L$ ).<sup>1</sup> This choice of physical dimensions and material for the wire simplifies the analysis of the results by satisfying the so-called ‘lumped-capacitance approximation’, which is characterized by a small Biot number ( $Bi=hD/k_{solid}<0.1$ , where  $h$  is the convection heat transfer coefficient and  $k_{solid}$  is the wire thermal conductivity). The measured room temperature resistance of this wire is  $R_{ho}=2.58\Omega$  (at  $295\text{ K}$ ). An adjustable resistor ( $R_p$ ) is used to balance the wire resistor  $R_h$  on the opposite leg of the bridge prior to heating the gold wire, as shown in Fig. 1, and we take  $R_1=2.29\Omega$ . The bridge is mounted on a  $3M^2$  Model 309 breadboard which facilitates changing bridge resistances, as needed. The wire is mounted across corner pads of a 40 pin dual-in-line package (DIP, Aries Model 40-6553-18 high-temperature test socket, max:  $523\text{ K}$ ). The DIP is incorporated directly into the bridge circuit (see Fig. 1).

A voltage pulse ( $V_{in}$ ) is imposed across the bridge of time duration  $\Delta t$  by an Agilent Model 8114A pulse generator. The output voltage ( $V_{out}$ ) is monitored by a LeCroy WaveRunner Model 44xi  $5\text{Gs/s}$  digital oscilloscope through a LeCroy Model AP033 differential amplifier, and the voltage is converted to wire resistance with the equations of bridge circuitry [12]. Data files are stored directly on the oscilloscope and later transferred to a personal computer for analysis. A calibration of wire resistance is used to obtain the average wire temperature during the power pulse.

A  $50\mu s$  pulse duration was selected and fixed throughout the study to provide a valid basis for comparing different fluids. The input voltage to the bridge was adjusted so that the average wire temperature in air was about  $560\text{ K}$  at the end of the  $50\mu s$  pulse. The choice of a  $50\mu s$  pulse mitigated thermal shock effects and potential damage associated with repeated thermal stressing of the wire. The pulse time was short enough to reduce the onset of

<sup>1</sup> Relevant properties of gold (at  $300\text{ K}$ ) are the following: solid density ( $19,300\text{ kg m}^{-3}$ ); specific heat ( $129\text{ J kg}^{-1}\text{ K}^{-1}$ ); thermal conductivity ( $318\text{ W m}^{-1}\text{ K}^{-1}$ ).

<sup>2</sup> Certain commercial equipment or materials are identified in this publication to specify adequately the experimental procedure. Such identification does not imply recommendation or endorsement by the National Institute of Standards and Technology, nor does it imply that the materials or equipment are necessarily the best available for this purpose.

natural convection and prevent vaporization of the liquids investigated (i.e., significant superheating can be sustained with pulse heating). The additional conditions of  $V_{in}=2.9\text{ V}$  and bridge resistance  $R_1=2.29\ \Omega$  ensure that the wire temperature is far from the melting point of gold (1337 K) at the end of the pulse, and an output signal of sufficient magnitude is obtained to differentiate between the examined fluids. The short-duration pulsing, which is accomplished with the aforementioned power electronics, is sufficient to assume a fast-response assessment.

The power ( $p$ ) generated by the wire at any instant of time ( $t$ ) is determined from the bridge equations as

$$p = V_{in}(t)^2 \frac{R_h(t)}{[R_1 + R_h(t)]^2} \quad (1)$$

Accounting for the change of  $R_h$  with heating, the power input during the  $50\ \mu\text{s}$  pulse duration is between 0.7 and 0.8 W.

Before a new fluid is examined, the wire is carefully cleaned by gentle flushing with methanol and baking in an oven for several hours. Power pulses are then imposed on the wire in air (parameters specified above) to ensure that the resulting evolution of temperature and average wire temperature after  $50\ \mu\text{s}$  is repeatable in air (achieving a maximum wire temperature  $T_{max}=560\text{ K}$ ), as noted above. Small systematic variations of input power are sometimes required to adjust the bridge impedance and  $V_{in}$  to achieve  $T_{max}$  for air. This effect may be due to annealing of the wire associated with repeated pulsing, which varied slightly the electrical conditions. By adjusting the input electrical parameters, using the peak temperature achieved in air after  $50\ \mu\text{s}$ , it was possible to ensure a uniform starting condition for examining the thermal response of the wire with each new fluid.

Both liquids and gases were examined. Measurements with liquids were carried out by placing a drop of the test fluid on the DIP recess to submerge the wire, and then powering the wire typically at a frequency of 10 Hz. Gases were studied by placing the DIP with bonded gold wire into a 2.54 cm diameter test tube, which was closed with a two-holed rubber stopper to provide a sealed containment for the gas. Glass tubes, which passed through the stopper holes, were connected to ball valves to control the gas flow through the test tube. The bonded wires, which were connected to the DIP and attached to the bridge circuit, were threaded out of the test tube and sealed by press fitting the wires against the rubber stopper. Before each experiment, the gas was allowed to flow through the test tube for several minutes. The ball valves were then closed quickly to seal the gas inside the test tube after which the wire was heated in the aforementioned manner.

The variation of gold wire resistance with temperature was calibrated by slowly heating the wire in a Thermolyne Model 48000 computer-controlled furnace while monitoring the temperature of the DIP with an Omega Model KMQXL-010U-12 (250  $\mu\text{m}$  diameter, sheathed, ungrounded, K-type) thermocouple that was placed in the space between the DIP and test socket. The thermocouple time response, as specified by the manufacturer, was estimated to be 0.5 s in water (0.8 s in air) at a level of confidence of 95%. The expanded uncertainty for the temperature was estimated to be 7.9 K, including the Type B uncertainty of 3.1 K,<sup>3</sup> and represented the major source of uncertainty in the measurement. The resistance was monitored during heating by an HP Model 3457 digital multimeter interfaced with a personal computer. Control of the calibration process was accomplished with a LABView program. Room temperature was recorded with

<sup>3</sup> Estimation of the measurement uncertainty is determined from statistical analysis of a series of replicated measurements (referred to as Type A evaluation of uncertainty) and from means other than statistical analysis (e.g., manufacturer estimates, referred to as Type B evaluation of uncertainty) [13].

an HP Model 2804A quartz probe, which served as the standard for calibration of the thermocouple.

The resistance calibration of the gold wire with varying oven temperature exhibited a slight hysteresis between the heating and cooling (having lower resistance values) portions of the calibration, which was attributed to annealing of the wire while cycling the temperature. A linear fit to the data was then used to correlate the wire resistance with the oven temperature ( $T_{TC}$ ) as

$$R_h = a + bT_{TC} \quad (2)$$

where  $a = -0.00331\ \Omega$  and  $b = 0.00877\ \Omega/\text{K}$ . The wire temperature ( $T$ ) is determined from Eq. (6) of Ref. [12], using the wire resistance obtained from the calibration, and  $a = R_{ho}[1 - \theta_R T_\infty]$ ,  $b = R_{ho}\theta_R$ , and  $\theta_R = 0.003394$ .

In Table 1, the first seven fluids are fire suppressant-related agents. The remaining ten fluids are not suppressants but are nonetheless included for completeness in the context of a larger class of fluids. Solid aerosol powders, water mists, and particle/fluid suspensions fire suppressants are not considered here. Issues arise regarding the influence of a temperature gradient imposed by the hot wire (i.e., which could impart a thermophoretic force to the particles and artificially influence their proximity to the wire) on suspended particles.

### 3. Results and discussion

The experimental results are presented in terms of a simple lumped thermal model for heat transfer between the wire and fluid for constant internal generation of heat and fixed ambient temperature. Assuming constant properties, a constant input power, and single-phase heat transfer, it can be shown that at the end of a power pulse (of duration  $t_{max}$ ) [14]

$$\theta_{max} = \frac{\bar{p}}{\xi} (1 - e^{-\xi \tau_{max}}) \quad (3)$$

where

$$\theta_{max} = \frac{T_{max} - T_\infty}{T_\infty}, \quad \tau_{max} = \frac{\alpha_{solid} t_{max}}{D^2}, \quad \bar{p} = \frac{4p}{\pi k_{solid} L T_\infty},$$

$$\text{and } \xi = 4Nu \frac{k_{fluid}}{k_{solid}}. \quad (4)$$

where  $\theta_{max}$  is the nondimensional temperature difference at  $T_{max}$ ,  $T_\infty$  is the ambient fluid temperature,  $\tau_{max}$  is the nondimensional time at  $t_{max}$ ,  $\alpha$  is the thermal diffusivity,  $k$  is the thermal conductivity, and  $\bar{p}$  is the nondimensional power ( $p$ ). The variable  $\xi$  is important for assessing the thermal behavior of a fire suppressant since it depends on the heat transfer mechanism between the fluid and wire (through the Nusselt number,  $Nu$ ), and the ratio of fluid-to-solid thermal conductivity. The Nusselt number is altered according to whether the flow between the fluid and wire is laminar or turbulent, and mechanism of heat transfer is conductive or convective.

Table 1 presents values of  $\xi$  along with the quenching effectiveness ( $QE$ ), which is defined as the ratio of the maximum temperature relative to that of water at the end of the  $50\ \mu\text{s}$  pulse,  $QE \equiv \theta_{max}/\theta_{max,ref}$  where

$$\theta_{max,ref} = \frac{T_{max,ref} - T_\infty}{T_\infty}$$

and  $T_{max,ref}$  (at  $50\ \mu\text{s}$ ) = 339 K for  $T_\infty \approx 295\text{ K}$  (the laboratory environment) as measured for water. The reported uncertainties for  $QE$  and  $\xi$  are based on relatively small Type A uncertainties for the electronic components, which supports the small variability in the measurements observed for each individual fluid (see the values of  $R_h$  in Table 1), and enabling us to confidently list the

**Table 1**Computed values of quenching effectiveness (*QE*) and other relevant fluid properties (fluids in bold are known fire suppressants and in italics are gases).

	Chemical name	Chemical formula	$R_{h,max}$ ( $\Omega$ )	$QE [= \theta_{max}/\theta_{max,ref}]$	$k/k_{ref}$	$\xi$
1	<b>water</b>	H <sub>2</sub> O	$2.97 \pm 4 \times 10^{-3}$	$1.00 \pm 4 \times 10^{-2}$	1.00 <sup>a</sup>	$2.50 \times 10^{-2} \pm 7 \times 10^{-4}$
2	<b>60% potassium acetate/water</b>	KC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> /H <sub>2</sub> O	$3.09 \pm 4 \times 10^{-3}$	$1.31 \pm 5 \times 10^{-2}$	0.80 <sup>b</sup>	$1.90 \times 10^{-2} \pm 4 \times 10^{-4}$
3	methanol	CH <sub>4</sub> O	$3.45 \pm 5 \times 10^{-3}$	$2.25 \pm 7 \times 10^{-2}$	0.30 <sup>c</sup>	$1.01 \times 10^{-2} \pm 1 \times 10^{-4}$
4	ethanol	C <sub>2</sub> H <sub>6</sub> O	$3.53 \pm 5 \times 10^{-3}$	$2.46 \pm 8 \times 10^{-2}$	0.24 <sup>c</sup>	$8.90 \times 10^{-3} \pm 1 \times 10^{-4}$
5	1-pentanol	C <sub>5</sub> H <sub>12</sub> O	$3.56 \pm 5 \times 10^{-3}$	$2.55 \pm 8 \times 10^{-2}$	0.22 <sup>c</sup>	$8.50 \times 10^{-3} \pm 1 \times 10^{-4}$
6	1-propanol	C <sub>3</sub> H <sub>8</sub> O	$3.58 \pm 5 \times 10^{-3}$	$2.58 \pm 8 \times 10^{-2}$	0.23 <sup>c</sup>	$8.36 \times 10^{-3} \pm 1 \times 10^{-4}$
7	n-decane	C <sub>10</sub> H <sub>22</sub>	$3.74 \pm 6 \times 10^{-3}$	$3.00 \pm 9 \times 10^{-2}$	0.18 <sup>c</sup>	$6.50 \times 10^{-3} \pm 9 \times 10^{-5}$
8	n-nonane	C <sub>9</sub> H <sub>20</sub>	$3.76 \pm 6 \times 10^{-3}$	$3.06 \pm 9 \times 10^{-2}$	0.17 <sup>c</sup>	$6.30 \times 10^{-3} \pm 8 \times 10^{-5}$
9	n-hexane	C <sub>6</sub> H <sub>14</sub>	$3.80 \pm 6 \times 10^{-3}$	$3.17 \pm 9 \times 10^{-2}$	0.15 <sup>c</sup>	$5.90 \times 10^{-3} \pm 8 \times 10^{-5}$
10	<b>1-methoxyheptafluoropropane (HFE7000)<sup>d</sup></b>	C <sub>4</sub> F <sub>7</sub> OH <sub>3</sub>	$4.25 \pm 7 \times 10^{-3}$	$4.32 \pm 1 \times 10^{-1}$	0.09 <sup>a</sup>	$2.76 \times 10^{-3} \pm 4 \times 10^{-5}$
11	<b>methoxy-nonfluorobutane (HFE7100)<sup>d</sup></b>	C <sub>5</sub> F <sub>9</sub> OH <sub>3</sub>	$4.26 \pm 7 \times 10^{-3}$	$4.36 \pm 1 \times 10^{-1}$	0.08 <sup>a</sup>	$2.60 \times 10^{-3} \pm 4 \times 10^{-5}$
12	<b>perfluoro-n-hexane (FC-72)<sup>e</sup></b>	C <sub>6</sub> F <sub>14</sub>	$4.27 \pm 7 \times 10^{-3}$	$4.38 \pm 1 \times 10^{-1}$	0.09 <sup>f</sup>	$2.52 \times 10^{-3} \pm 4 \times 10^{-5}$
13	<i>helium</i>	He	$4.62 \pm 7 \times 10^{-3}$	$5.28 \pm 2 \times 10^{-1}$	0.30 <sup>g</sup>	$1.35 \times 10^{-3} \pm 3 \times 10^{-5}$
14	<i>air</i>	–	$4.84 \pm 8 \times 10^{-3}$	$5.86 \pm 2 \times 10^{-1}$	0.06 <sup>g</sup>	$4.00 \times 10^{-4} \pm 1 \times 10^{-5}$
15	<b>1,1,1,2,2-pentafluoroethane (HFC-125)</b>	C <sub>2</sub> HF <sub>5</sub>	$4.84 \pm 8 \times 10^{-3}$	$5.86 \pm 2 \times 10^{-1}$	0.05 <sup>h</sup>	$4.00 \times 10^{-4} \pm 1 \times 10^{-5}$
16	<b>bromotrifluoromethane (Halon 1301)</b>	CF <sub>3</sub> Br	$4.91 \pm 8 \times 10^{-3}$	$6.03 \pm 2 \times 10^{-1}$	0.03 <sup>i</sup>	$1.00 \times 10^{-4} \pm 9 \times 10^{-6}$

The term  $\xi$  was determined iteratively from Eq. (3). For water as the reference,  $\theta_{max,ref} = 0.149 \pm 4 \times 10^{-3}$  and  $k_{ref} = 0.638 \text{ W m}^{-1} \text{ K}^{-1}$ . The reported uncertainties are calculated as  $k_c u_c$ , where  $k_c$  is the coverage factor ( $=2$ ) and  $u_c$  is the combined standard uncertainty.

<sup>a</sup> C. Presser, C.T. Avedisian, Atomization and Sprays 16(6) (2006) 627–656.

<sup>b</sup> US Patent No. 6893582, May 17, 2005.

<sup>c</sup> D.R. Lide (Ed.), CRC Handbook of Chemistry and Physics, 87th ed., Taylor & Francis Group, Florida, 2006.

<sup>d</sup> 3M™ Novec™ Engineered Fluid.

<sup>e</sup> 3M™ Fluorinert™ Electronic Liquid.

<sup>f</sup> 3M™ Fluorinert™ Electronic Liquid FC-72 Product Information Sheet, 98-0212-2308-0 (HB), 3M Specialty Materials, St. Paul, MN, 2000.

<sup>g</sup> <http://en.wikipedia.org/wiki/Wikipedia>.

<sup>h</sup> R.A. Perkins, M.L. Huber, Journal of Chemical and Engineering Data 51 (2006) 898–904.

<sup>i</sup> C. Baroncini, P. Di Filippo, G. Latini, M. Pacetti, International Journal of Thermophysics 2(1) (1981) 21–38.

values of *QE* in the ascending order given in Table 1. The major component of uncertainty lies with the thermocouple Type B values (as provided by the manufacturer) that was used for the calibration, however, since the reported temperatures were based on the linear fit through the calibration data, the relative effect on the results given in Table 1 will be negligible. The term *QE* is a measure of a fluid's ability to suppress the temperature change of a surface suddenly to an ambience at a different temperature. It may be larger or smaller than unity depending on the value of  $\theta_{max,ref}$ . Among a class of fluids, the most effective suppressant from a thermal standpoint has the smallest *QE* and highest  $\xi$ . On this basis, water is the most thermally effective agent listed in Table 1.

Comparing the fluids in Table 1, all of the liquids have a smaller value of *QE* than the gases (given in italics). Among the gases in Table 1, helium is the most effective (having the smallest value) principally because of its relatively high thermal conductivity. The value of *QE* for the hydrofluoroethane C<sub>4</sub>F<sub>7</sub>OH<sub>3</sub> is somewhat smaller than C<sub>5</sub>F<sub>9</sub>OH<sub>3</sub> and, thus, C<sub>4</sub>F<sub>7</sub>OH<sub>3</sub> may be slightly preferable to C<sub>5</sub>F<sub>9</sub>OH<sub>3</sub> for fire suppression. The perfluorocarbon C<sub>6</sub>F<sub>14</sub> has a slightly larger value of *QE* than the two hydrofluoroethanes. The alcohols (not relevant to fire suppression but interesting nonetheless to view their ranking) have higher values of *QE* than the alkanes.

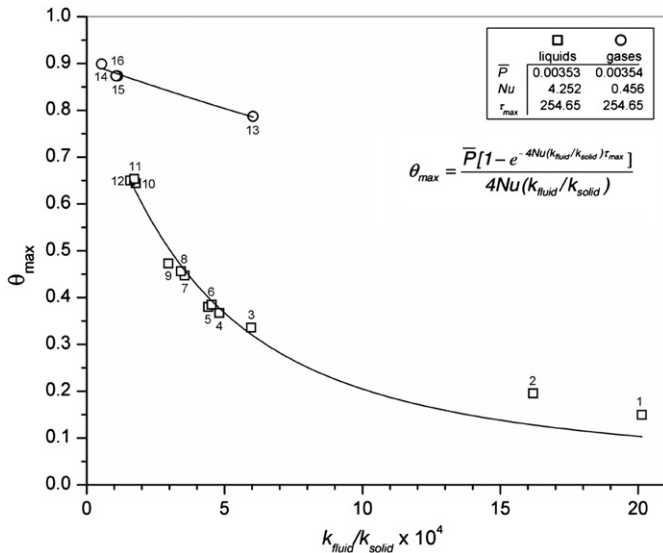
Although gases appear to be less thermally effective than liquids, it is well known that certain gases are extremely effective fire suppressants. The case of CF<sub>3</sub>Br is especially interesting in the context of the present study. It has the largest value of *QE* (smallest value of  $\xi$ ) compared to all of the fluids examined and, so, would be considered a less thermally effective suppressant than the other fluids listed in Table 1. However, CF<sub>3</sub>Br is known to be an extremely effective suppressant. The reason attributed to chemical effects, in which Br recombines catalytically H or OH radicals to reduce the radical pool and lead to flame extinction [1,2]. This example illustrates that thermal nor chemical

considerations alone dictate the effectiveness of a fluid for fire suppression (indeed, flammable liquids with no relevance to fire suppression have higher values of *QE* (lower values of  $\xi$ ) than fluids which are known to be effective suppressants, as indicated in Table 1. It will certainly be true that fluids, which are effective both chemically and thermally (i.e., having lower values of *QE* and higher values of  $\xi$ ) will be preferable to fluids with the opposite thermal characteristics.

Eqs. (3) and (4) suggest that  $\theta_{max}$  should correlate with  $k_{fluid}/k_{solid}$  for a given value of *Nu* and  $\bar{p}$ . Fig. 2 presents the variation of  $\theta_{max}$  with the ratio of fluid-to-solid thermal conductivity for the fluids listed in Table 1. The correlations for gases and liquids are segregated for a given  $\bar{p}$  – a consequence of the large difference in liquid and gas properties, which produce  $Nu_{liquid} > Nu_{gas}$ . A regression analysis was used to fit the data, using *Nu* and  $\bar{p}$  as parameters. The results indicate that the input power is almost the same for the gases and liquids examined (i.e.,  $\bar{p}_{gas} = 0.00354$  and  $\bar{p}_{liquid} = 0.00353$ ), while values of  $Nu_{gas} = 0.456$  and  $Nu_{liquid} = 4.252$  provide the best match to the measurements. The gas Nusselt number is remarkably close to the conduction limit,  $Nu = 0.36$  [14], while the value of  $Nu_{liquid}$  suggests that heating in the liquids may have induced some level of convection. The lines drawn in Fig. 2 are based on Eq. (3), using the stated values for the nondimensional power and Nusselt number. The close agreement of Eq. (3) with the data shown in Fig. 2 suggests that the simplified model represents reasonably well the experimental measurements.

#### 4. Summary

A screening concept, based on the transient hot-wire technique, is presented to evaluate the thermal potential of a liquid or gaseous fire suppressant to restrain the temperature change associated with interaction of a cold fluid with a comparatively



**Fig. 2.** Variation of  $\theta_{max}$  with ratio of fluid-to-solid thermal conductivity. Parameters are  $\bar{P}$  and  $Nu$ . Inset shows regression values that best represent the measurements in Table 1. Numbers correspond to those given for the fluids listed in Table 1.

hot (or burning) surface. A nondimensional parameter based on inhibiting either the temporal change in wire surface temperature or the maximum temperature reached at the end of the power pulse is used to rank different chemical agents. The liquids as a group appear to be better thermally effective suppressants than the gases. Water is considered the 'best' thermally effective suppressant (surface coolant) from among the fluids examined. A lumped thermal system was found to reproduce the measurements with appropriate values of the Nusselt number.

## References

- [1] P. Papas, J.W. Fleming, R.S. Sheinson, Extinction of non-premixed methane- and propane-air counterflow flames inhibited with  $CF_4$ ,  $CF_3H$  and  $CF_3Br$ , in: 26th Symposium (International) on Combustion, Combustion Institute, Pittsburgh, PA, 1996, pp. 1405–1411.
- [2] G.T. Linteris, L. Truett, Inhibition of premixed methane-air flames by fluoromethanes, *Combust Flame* 105 (1996) 15–27.
- [3] A. Hamins, et al., Suppression of engine nacelle fires, in: R.G. Gann (Ed.), *Fire Suppression System Performance of Alternative Agents in Aircraft Engine and Dry Bay Laboratory Simulations*, NIST SP 890, vol. II, National Institute of Standards and Technology, Gaithersburg, MD, USA, 1995, pp. 1–199 (Chapter 9).
- [4] G.T. Linteris, V.R. Katta, F. Takahashi, Experimental and numerical evaluation of metallic compounds for suppressing cup-burner flames, *Combust Flame* 138 (1–2) (2004) 78–96.
- [5] J.C. Yang, M.K. Donnelly, N.C. Prive, W.L. Grosshandler, An apparatus for screening fire suppression efficiency of dispersed liquid agents, *Fire Safety Journal* 36 (2001) 55–72.
- [6] G.T. Linteris, F. Takahashi, V.R. Katta, Cup-burner flame extinguishment by  $CF_3Br$  and  $Br_2$ , *Combust Flame* 149 (1–2) (2007) 91–103.
- [7] A. Hamins, C. Presser, L. Melton, Suppression of a baffle stabilized spray flame by halogenated agents, *Proceedings of the Combustion Institute*, vol. 26, The Combustion Institute, Pittsburgh, PA, 1996, pp. 1413–1420.
- [8] W.M. Pitts, J.C. Yang, R.A. Bryant, L.G. Blevins, M.L. Huber, Characterization and identification of super-effective thermal fire extinguishing agents: Final Report, NGP Project 4C/1/890, NIST Technical Note 1440, National Institute of Standards and Technology, Gaithersburg, MD, USA, July 2006.
- [9] A.M. Lentati, H.K. Chelliah, Physical, thermal, and chemical effects of fine-water droplets in extinguishing counterflow diffusion flames, in: 27th Symposium (International) on Combustion, Combustion Institute, Pittsburgh, PA, USA, 1998, pp. 2839–2846.
- [10] A.K. Lazzarini, R.H. Krauss, H.K. Chelliah, G.T. Linteris, Extinction conditions of nonpremixed flames with fine droplets of water and water–NaOH solutions, in: 28th Symposium (International) on Combustion, Combustion Institute, Pittsburgh, PA, USA, 2000, pp. 2939–2945.
- [11] C.T. Avedisian, W.S. Osborne, F.D. McLeod, C.M. Curley, Measuring bubble nucleation temperature on the surface of a rapidly heated thermal ink-jet heater immersed in a pool of water, *Proc. R. Soc. London A* 455 (1999) 3875–3899.
- [12] O.C. Thomas, R.E. Cavicchi, M.J. Tarlov, Effect of surface wettability on fast transient microboiling behavior, *Langmuir* 19 (15) (2003) 6168–6177.
- [13] B.N. Taylor, C.E. Kuyatt, Guidelines for evaluating and expressing the uncertainty of NIST measurement results, NIST Technical Note 1297, National Institute of Standards and Technology, Gaithersburg, 1994.
- [14] F.P. Incropera, D.P. DeWitt, *Introduction to Heat Transfer*, 4th ed, Wiley, New York, 2001, p. 248, 516.