

Electrochemistry

I. Balancing redox equations

- half reaction method

II. Galvanic (Voltaic) cells

- anode and cathode
- cell diagram
- cell notation
- cell potential

III. Standard reduction potentials

- standard emf (E°_{cell})
- standard hydrogen electrode (SHE)

IV. Relating E° to ΔG° and K

V. Dependence of cell potential on concentration

- the Nernst equation

VI. Batteries

VII. Corrosion

VIII. Electrolysis